



## Cambridge O Level

CANDIDATE  
NAME

CENTRE  
NUMBER

--	--	--	--	--

CANDIDATE  
NUMBER

--	--	--	--



**CHEMISTRY**

**5070/21**

Paper 2 Theory

**October/November 2022**

**1 hour 30 minutes**

You must answer on the question paper.

No additional materials are needed.

### INSTRUCTIONS

- Section A: answer **all** questions.
- Section B: answer **three** questions.
- Use a black or dark blue pen. You may use an HB pencil for any diagrams or graphs.
- Write your name, centre number and candidate number in the boxes at the top of the page.
- Write your answer to each question in the space provided.
- Do **not** use an erasable pen or correction fluid.
- Do **not** write on any bar codes.
- You may use a calculator.
- You should show all your working and use appropriate units.

### INFORMATION

- The total mark for this paper is 75.
- The number of marks for each question or part question is shown in brackets [ ].
- The Periodic Table is printed in the question paper.

This document has **20** pages. Any blank pages are indicated.

## Section A

Answer **all** the questions in this section in the spaces provided.

The total mark for this section is 45.

1 The diagram shows part of the Periodic Table.

I	II									III	IV	V	VI	VII	VIII
												N	O		
Na										Al	Si			Cl	
K	Ca				Cr		Fe		Ni		Zn				
	Sr									Ag					
	Ba														

Answer the following questions using only the symbols of the elements in the diagram.

Each symbol may be used once, more than once or not at all.

Give the symbol of the element that:

(a) is a catalyst in the hydrogenation of alkenes

..... [1]

(b) is extracted by electrolysis of its oxide dissolved in cryolite

..... [1]

(c) is a metal below copper in the reactivity series

..... [1]

(d) is a gas used in welding

..... [1]

(e) forms an ion with a charge of  $-1$ .

..... [1]

[Total: 5]

2 Carbon dioxide is a gas at room temperature.

(a) (i) Draw a dot-and-cross diagram to show the arrangement of electrons in a carbon dioxide molecule.

Show only the outer shell electrons.

[2]

(ii) Describe the motion and separation of the particles in a gas.

motion .....

separation .....

[2]

(b) The main processes in the carbon cycle are combustion, respiration and photosynthesis.

(i) Name the products of photosynthesis.

..... and ..... [1]

(ii) Some of the reactions in photosynthesis involve enzymes.

State the meaning of the term *enzyme*.

.....

..... [1]

(c) Carbon dioxide is formed when hydrocarbons are completely combusted.

(i) Name the other product of the complete combustion of hydrocarbons.

..... [1]

(ii) The paraffin (kerosene) fraction from the fractional distillation of petroleum (crude oil) contains hydrocarbons.

State **one** use of the paraffin (kerosene) fraction.

..... [1]

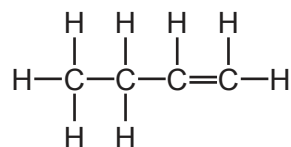
[Total: 8]

3 The alkenes are a homologous series of hydrocarbons.

(a) Give the general formula for the alkenes.

..... [1]

(b) The structure of an alkene is shown.



(i) Name this alkene.

..... [1]

(ii) Explain how this structure shows that alkenes are:

unsaturated .....

.....

hydrocarbons. ....

.....

[2]

(c) Alkenes are produced by cracking some fractions obtained from the fractional distillation of petroleum (crude oil).

State the meaning of the term *cracking*.

.....

..... [2]

(d) Carbon monoxide is formed when alkenes undergo incomplete combustion.

State **one** effect of carbon monoxide on health.

..... [1]

- (e) Alkenes react with bromine to form compounds containing carbon, hydrogen and bromine.
- (i) A compound contains 22.2% carbon, 3.70% hydrogen and 74.1% bromine by mass.

Calculate the empirical formula of this compound.

empirical formula ..... [2]

- (ii) A different compound of carbon, hydrogen and bromine has the empirical formula  $C_3H_2Br$ .

The relative molecular mass of this compound is 236.

Deduce the molecular formula of this compound.

molecular formula ..... [1]

[Total: 10]

4 This question is about ammonia and ammonium salts.

(a) Ammonia is manufactured by the Haber process.

(i) Name the catalyst used in the Haber process.

..... [1]

(ii) Explain how a catalyst increases the rate of reaction.

.....  
 ..... [1]

(b) Ammonium phosphate,  $(\text{NH}_4)_3\text{PO}_4$ , is a fertiliser.

Calculate the percentage by mass of nitrogen in ammonium phosphate.

Give your answer to **three** significant figures.

percentage by mass = ..... [3]

(c) Fertilisers can make soil more acidic.

(i) State the name of a compound used to decrease the acidity of soil.

..... [1]

(ii) Explain how this compound decreases the acidity of soil.

..... [1]

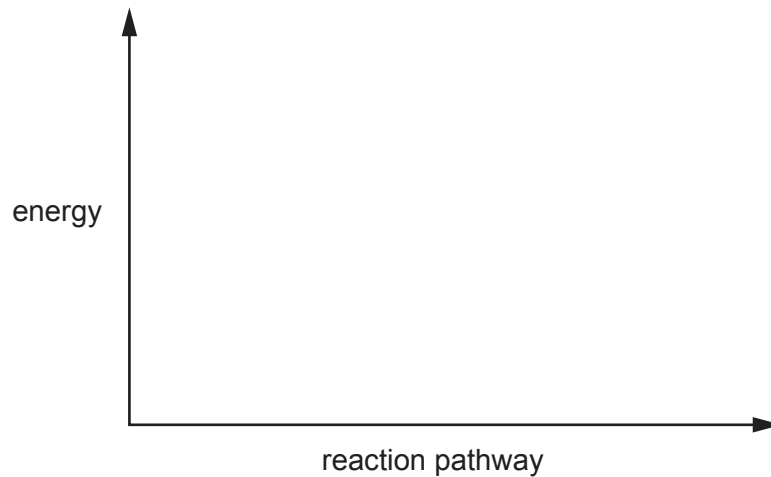
(d) Ammonia is formed by the reduction of nitrogen(I) oxide,  $\text{N}_2\text{O}$ , with hydrogen.

The reaction is exothermic.



Complete and label the energy profile diagram for this reaction to include:

- the reactants and products
- the enthalpy change of the reaction.



[2]

(e) Copper(II) oxide,  $\text{CuO}$ , reacts with ammonia.

The products are copper, nitrogen and a liquid which turns blue cobalt(II) chloride paper pink.

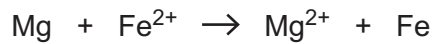
Construct the equation for this reaction.

..... [2]

[Total: 11]

5 This question is about metals and metal compounds.

(a) Magnesium reacts with aqueous iron(II) ions.



Explain why this reaction involves **both** oxidation and reduction.

Use the equation and ideas about electron transfer in your answer.

.....  
 .....  
 .....  
 ..... [2]

(b) Explain why iron conducts electricity.

..... [1]

(c) Explain why magnesium is extracted by electrolysis and not by reduction with carbon.

..... [1]

(d) Describe a chemical test to distinguish between aqueous iron(II) ions and aqueous iron(III) ions.

test .....

observations with aqueous iron(II) ions

.....

observations with aqueous iron(III) ions

..... [3]

(e) (i) Explain why molten magnesium chloride conducts electricity.

..... [1]

(ii) Predict the products formed at the anode and the cathode when molten magnesium chloride is electrolysed.

anode .....

cathode .....

[2]



(f) Food containers can be made from aluminium.

State **one** property of aluminium that makes it suitable for use as a food container.

..... [1]

[Total: 11]

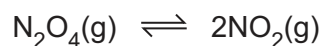
## Section B

Answer **three** questions from this section in the spaces provided.

The total mark for this section is 30.

6 This question is about compounds of nitrogen.

- (a) The equation represents the equilibrium between  $\text{N}_2\text{O}_4$  and  $\text{NO}_2$  at a high temperature in a closed container.



- (i) Predict what happens to the position of equilibrium when the pressure is decreased.

Explain your answer.

prediction .....

explanation .....

.....

.....

[2]

- (ii) The table shows the concentration of  $\text{NO}_2$  in the closed container at three different temperatures.

temperature in °C	concentration of $\text{NO}_2$ in mol/dm <sup>3</sup>
100	0.04
150	0.40
200	4.00

State what this information shows about the enthalpy change of the forward reaction.

Explain your answer.

enthalpy change .....

explanation .....

.....

.....

[2]

(b) Nitrogen dioxide,  $\text{NO}_2$ , is made by heating lead(II) nitrate,  $\text{Pb}(\text{NO}_3)_2$ .  
The other products are lead(II) oxide and a gas which relights a glowing splint.

(i) Construct the equation for this reaction.

..... [2]

(ii) Nitrogen dioxide contributes to acid rain.

State **one** effect of acid rain on organisms.

..... [1]

(c) Nitric acid,  $\text{HNO}_3$ , is a strong acid.

(i) State the meaning of the term *strong* in strong acid.

.....  
..... [1]

(ii) Suggest a pH value for a concentrated solution of a strong acid.

..... [1]

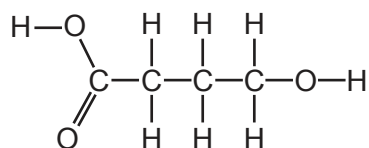
(iii) Complete the ionic equation for the reaction of an acid with an alkali.



[Total: 10]

7 This question is about alcohols and polymers.

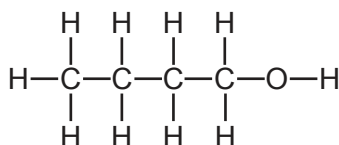
(a) The structure of an organic compound is shown.



Deduce the molecular formula of this compound.

..... [1]

(b) An isomer of butanol has the structure shown.



(i) State the meaning of the term *isomerism*.

.....  
 ..... [1]

(ii) Draw the structure of a different isomer of butanol.

[1]

(iii) Butanol reacts with ethanoic acid,  $\text{CH}_3\text{COOH}$ , to form an ester.

Name this ester.

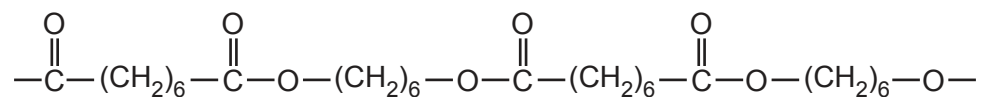
Draw the structure of this ester. Show all of the atoms and all of the bonds.

name of ester .....

structure of ester

[2]

(c) The partial structure of a polyester is shown.



Draw the structures of the two monomers used to make this polyester.

[2]

(d) *Terylene* is a polyester.

(i) State **one** use of *Terylene*.

..... [1]

(ii) Name a naturally occurring molecule which has the same ester linkage as *Terylene*.

..... [1]

(e) Starch is a polymer which can be hydrolysed.

State the meaning of the term *hydrolysis*.

..... [1]

[Total: 10]

8 This question is about metals and metal compounds.

(a) Brass is an alloy of copper and zinc.

Brass is less malleable than either copper or zinc.

Explain, with reference to the structure of metals, why brass is less malleable than either copper or zinc.

You may include a labelled diagram in your answer.

.....

.....

.....

..... [3]

(b) The table shows the reactivity of four metals with cold water and with steam.

metal	reactivity with cold water	reactivity with steam
chromium	none	slow
mercury	none	none
potassium	very fast	explosive
uranium	very slow	slow

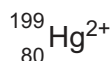
Put the four metals in order of increasing reactivity.

least reactive  $\xrightarrow{\hspace{15em}}$  most reactive

[1]

(c) The full symbol of an ion of mercury is shown.



Deduce the number of electrons and neutrons in this ion.

number of electrons .....

number of neutrons .....

[2]

(d) The formula of chromium(III) sulfate crystals is  $\text{Cr}_2(\text{SO}_4)_3 \cdot 18\text{H}_2\text{O}$ .

State the name given to the water present in  $\text{Cr}_2(\text{SO}_4)_3 \cdot 18\text{H}_2\text{O}$ .

..... [1]

(e) (i) Aqueous bromine reacts with aqueous potassium iodide.

The products of the reaction are aqueous iodine and aqueous potassium bromide.

Construct the ionic equation, including state symbols, for this reaction.

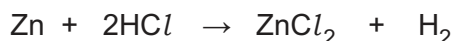
..... [2]

(ii) Explain, in terms of the reactivity of the halogens, why aqueous iodine does **not** react with aqueous potassium bromide.

..... [1]

[Total: 10]

- 9 (a) Zinc powder reacts with dilute hydrochloric acid.



When  $20.0\text{ cm}^3$  of dilute hydrochloric acid is added to excess zinc, the volume of hydrogen gas produced at room temperature and pressure is  $60.0\text{ cm}^3$ .

- (i) Calculate the concentration, in  $\text{mol/dm}^3$ , of the dilute hydrochloric acid.

concentration .....  $\text{mol/dm}^3$  [3]

- (ii) The reaction is repeated using large pieces of zinc instead of zinc powder. All other conditions stay the same.

Describe how the rate of reaction changes.

Explain your answer using ideas about collisions between particles.

.....  
 .....  
 ..... [2]

- (b) Dilute hydrochloric acid is electrolysed.

Construct an ionic equation for the reaction at the cathode.

..... [1]

- (c) Zinc is used to prevent iron from rusting.

- (i) State the essential conditions needed for rusting.

..... [1]

- (ii) Explain how zinc prevents iron from rusting by the method of sacrificial protection.

.....  
 ..... [2]

- (iii) Give **one** example of the use of sacrificial protection.

..... [1]

[Total: 10]







**BLANK PAGE**

---

Permission to reproduce items where third-party owned material protected by copyright is included has been sought and cleared where possible. Every reasonable effort has been made by the publisher (UCLES) to trace copyright holders, but if any items requiring clearance have unwittingly been included, the publisher will be pleased to make amends at the earliest possible opportunity.

To avoid the issue of disclosure of answer-related information to candidates, all copyright acknowledgements are reproduced online in the Cambridge Assessment International Education Copyright Acknowledgements Booklet. This is produced for each series of examinations and is freely available to download at [www.cambridgeinternational.org](http://www.cambridgeinternational.org) after the live examination series.

Cambridge Assessment International Education is part of Cambridge Assessment. Cambridge Assessment is the brand name of the University of Cambridge Local Examinations Syndicate (UCLES), which is a department of the University of Cambridge.

## The Periodic Table of Elements

		Group															
I	II											III	IV	V	VI	VII	VIII
3 <b>Li</b> lithium 7	4 <b>Be</b> beryllium 9	<b>Key</b> atomic number atomic symbol name relative atomic mass										5 <b>B</b> boron 11	6 <b>C</b> carbon 12	7 <b>N</b> nitrogen 14	8 <b>O</b> oxygen 16	9 <b>F</b> fluorine 19	10 <b>Ne</b> neon 20
11 <b>Na</b> sodium 23	12 <b>Mg</b> magnesium 24											13 <b>Al</b> aluminium 27	14 <b>Si</b> silicon 28	15 <b>P</b> phosphorus 31	16 <b>S</b> sulfur 32	17 <b>Cl</b> chlorine 35.5	18 <b>Ar</b> argon 40
19 <b>K</b> potassium 39	20 <b>Ca</b> calcium 40	21 <b>Sc</b> scandium 45	22 <b>Ti</b> titanium 48	23 <b>V</b> vanadium 51	24 <b>Cr</b> chromium 52	25 <b>Mn</b> manganese 55	26 <b>Fe</b> iron 56	27 <b>Co</b> cobalt 59	28 <b>Ni</b> nickel 59	29 <b>Cu</b> copper 64	30 <b>Zn</b> zinc 65	31 <b>Ga</b> gallium 70	32 <b>Ge</b> germanium 73	33 <b>As</b> arsenic 75	34 <b>Se</b> selenium 79	35 <b>Br</b> bromine 80	36 <b>Kr</b> krypton 84
37 <b>Rb</b> rubidium 85	38 <b>Sr</b> strontium 88	39 <b>Y</b> yttrium 89	40 <b>Zr</b> zirconium 91	41 <b>Nb</b> niobium 93	42 <b>Mo</b> molybdenum 96	43 <b>Tc</b> technetium —	44 <b>Ru</b> ruthenium 101	45 <b>Rh</b> rhodium 103	46 <b>Pd</b> palladium 106	47 <b>Ag</b> silver 108	48 <b>Cd</b> cadmium 112	49 <b>In</b> indium 115	50 <b>Sn</b> tin 119	51 <b>Sb</b> antimony 122	52 <b>Te</b> tellurium 128	53 <b>I</b> iodine 127	54 <b>Xe</b> xenon 131
55 <b>Cs</b> caesium 133	56 <b>Ba</b> barium 137	57–71 lanthanoids	72 <b>Hf</b> hafnium 178	73 <b>Ta</b> tantalum 181	74 <b>W</b> tungsten 184	75 <b>Re</b> rhenium 186	76 <b>Os</b> osmium 190	77 <b>Ir</b> iridium 192	78 <b>Pt</b> platinum 195	79 <b>Au</b> gold 197	80 <b>Hg</b> mercury 201	81 <b>Tl</b> thallium 204	82 <b>Pb</b> lead 207	83 <b>Bi</b> bismuth 209	84 <b>Po</b> polonium —	85 <b>At</b> astatine —	86 <b>Rn</b> radon —
87 <b>Fr</b> francium —	88 <b>Ra</b> radium —	89–103 actinoids	104 <b>Rf</b> rutherfordium —	105 <b>Db</b> dubnium —	106 <b>Sg</b> seaborgium —	107 <b>Bh</b> bohrium —	108 <b>Hs</b> hassium —	109 <b>Mt</b> meitnerium —	110 <b>Ds</b> darmstadtium —	111 <b>Rg</b> roentgenium —	112 <b>Cn</b> copernicium —	114 <b>Fl</b> flerovium —	116 <b>Lv</b> livermorium —	—	—	—	—

lanthanoids	57 <b>La</b> lanthanum 139	58 <b>Ce</b> cerium 140	59 <b>Pr</b> praseodymium 141	60 <b>Nd</b> neodymium 144	61 <b>Pm</b> promethium —	62 <b>Sm</b> samarium 150	63 <b>Eu</b> europium 152	64 <b>Gd</b> gadolinium 157	65 <b>Tb</b> terbium 159	66 <b>Dy</b> dysprosium 163	67 <b>Ho</b> holmium 165	68 <b>Er</b> erbium 167	69 <b>Tm</b> thulium 169	70 <b>Yb</b> ytterbium 173	71 <b>Lu</b> lutetium 175
actinoids	89 <b>Ac</b> actinium —	90 <b>Th</b> thorium 232	91 <b>Pa</b> protactinium 231	92 <b>U</b> uranium 238	93 <b>Np</b> neptunium —	94 <b>Pu</b> plutonium —	95 <b>Am</b> americium —	96 <b>Cm</b> curium —	97 <b>Bk</b> berkelium —	98 <b>Cf</b> californium —	99 <b>Es</b> einsteinium —	100 <b>Fm</b> fermium —	101 <b>Md</b> mendelevium —	102 <b>No</b> nobelium —	103 <b>Lr</b> lawrencium —

The volume of one mole of any gas is 24 dm<sup>3</sup> at room temperature and pressure (r.t.p.).